

Name: _____

Homework from the book:

Ch 4 Multiple Choice: 2, 6, 12, 14, 16-19, 32-34, 42 Exercises 61

Study Guide : Multiple Choice: 2, 5, 7, 11, 14, Solved problems 1-1 and 4-2

Ch 10 Multiple Choice: 3, 21, 27, 28, 30-34, 36-39 Exercises 33

Take home quiz:

- 1) From the diagram, identify the transition state, the activation energy and the energy of reaction.
- 2) Which of the following will not increase the rate of a reaction.
(A) Increase temperature. (B) Decrease concentration. (C) Add a catalyst.
- 3) Which of the following pH values describes an acidic solution?
(A) 3.52 (B) 7 (C) 8.4 (D) All of these are acidic. (E) None of these are acidic.
- 4) When this equation is balanced, what are the coefficients?
$$\underline{\hspace{1cm}} \text{Ba}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{H}_2\text{S} \rightarrow \underline{\hspace{1cm}} \text{BaS} + \underline{\hspace{1cm}} \text{HNO}_3$$

a) 1,2,1,2 b) 1,1,2,2 c) 1,1,1,2 d) 1,1,1,1 e) 2,2,2,2
- 5) Aspen is at an elevation of 7192 feet above sea level, Jacksonville is approximately 6 feet above sea level. The atmospheric pressure in Aspen is _____ than in Jacksonville.
A) higher B) lower C) the same D) can't tell
- 6) How many moles are present in a gas at 77°C and 1.33 atm that occupies 32.3 L?
- 7) The gas in an aerosol can is at a pressure of 3.00 atm at 25 °C. What will the pressure be if the temperature of the can is raised to 545 °C? (Assume a rigid container $V_1=V_2$)

